



Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CHEMISTRY

0620/12

Paper 1 Multiple Choice

May/June 2014

45 Minutes

Additional Materials: Multiple Choice Answer Sheet
 Soft clean eraser
 Soft pencil (type B or HB is recommended)



READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

DO NOT WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 16.

Electronic calculators may be used.

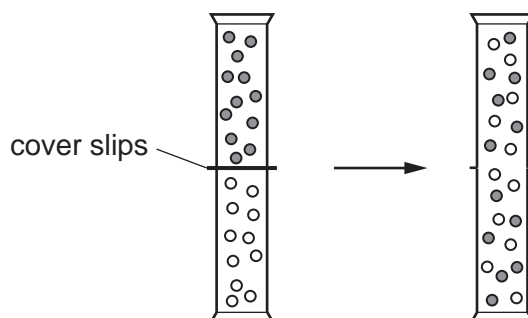
The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **16** printed pages.

2

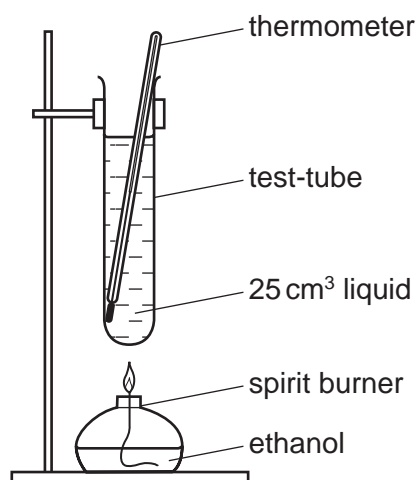
- 1 Two gas jars each contain a different gas. The gas jars are connected and the cover slips are removed.

The diagram shows what happens to the particles of the gases.



Which process has occurred?

- A chemical reaction
 - B condensation
 - C diffusion
 - D evaporation
- 2 A liquid is heated until it boils.



Which result shows that the liquid in the test-tube is pure water?

- A Condensation forms at the top of the test-tube.
- B Steam is produced.
- C The thermometer reads 100 °C.
- D There is nothing left behind in the test-tube.

3

3 Which two methods can be used to separate a salt from its solution in water?

- 1 crystallisation
- 2 decanting
- 3 distillation
- 4 filtration

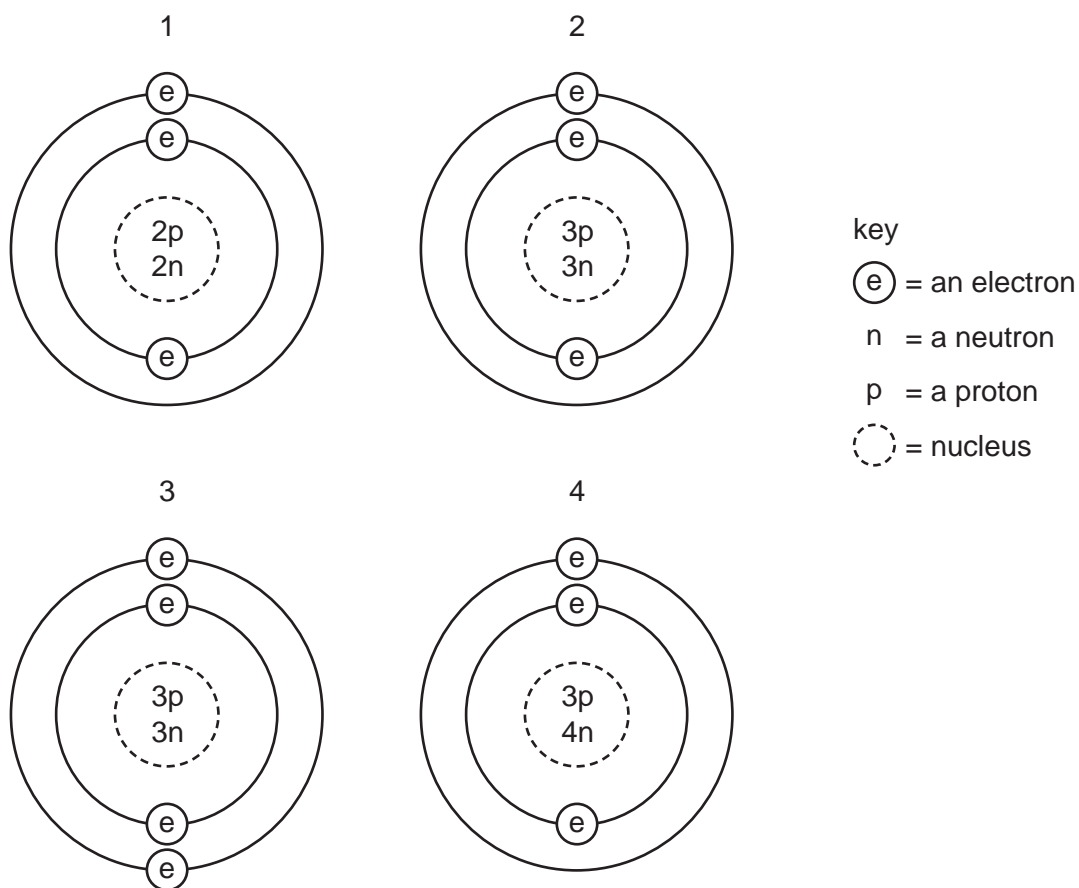
A 1 and 2 **B** 1 and 3 **C** 2 and 3 **D** 3 and 4

4 Which statements about a phosphorus atom, $^{31}_{15}\text{P}$, are correct?

- 1 The nucleon number is 16.
- 2 The number of outer electrons is 5.
- 3 The proton number is 15.

A 1, 2 and 3 **B** 1 and 2 only **C** 1 and 3 only **D** 2 and 3 only

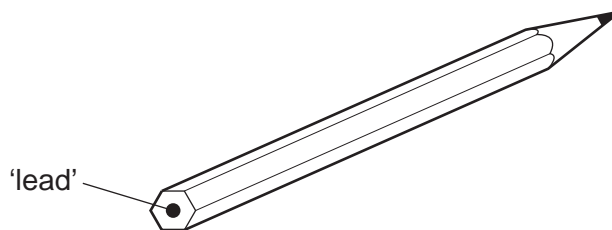
5 The diagrams show four particles.



Which two diagrams show **atoms** that are isotopes of each other?

A 1 and 2 **B** 1 and 3 **C** 2 and 3 **D** 2 and 4

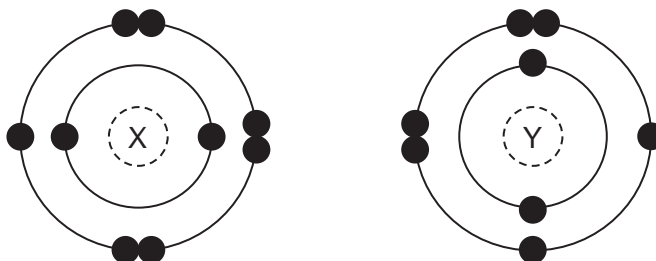
- 6 The 'lead' in a pencil is made of a mixture of graphite and clay.



When the percentage of graphite is increased, the pencil slides across the paper more easily.

Which statement explains this observation?

- A Graphite has a high melting point.
 - B Graphite is a form of carbon.
 - C Graphite is a lubricant.
 - D Graphite is a non-metal.
- 7 The electronic structures of two atoms, X and Y, are shown.



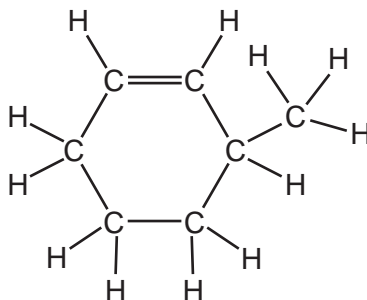
X and Y combine together to form a compound.

What is the type of bonding in the compound and what is the formula of the compound?

	type of bonding	formula
A	covalent	X_2Y
B	covalent	XY_2
C	ionic	XY_2
D	ionic	X_2Y

5

- 8 The structure of an organic compound, X, is shown.



What is the molecular formula of X?

- A C_6H_9 B C_6H_{12} C C_7H_{12} D C_7H_{14}
- 9 What is the relative molecular mass, M_r , of nitrogen dioxide?
- A 15 B 23 C 30 D 46
- 10 Electrical cables are made from either1....., because it is a very good conductor of electricity, or from.....2....., because it has a low density. Overhead cables have a3..... core in order to give the cable strength.

Which words correctly complete gaps 1, 2 and 3?

	1	2	3
A	aluminium	copper	magnesium
B	copper	aluminium	magnesium
C	copper	aluminium	steel
D	magnesium	copper	steel

- 11 What will be produced at the anode and at the cathode, if molten potassium chloride is electrolysed?

	anode (+)	cathode (-)
A	chlorine	hydrogen
B	chlorine	potassium
C	hydrogen	chlorine
D	potassium	chlorine

12 Solutions of two chemicals are mixed.

A reaction occurs and the temperature change is measured.

Which statement is correct?

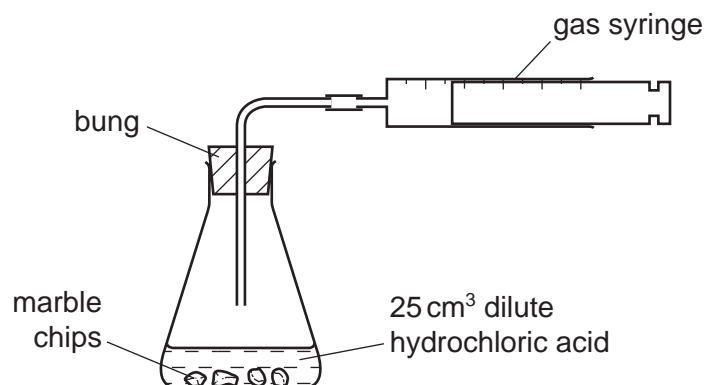
- A If the reaction is endothermic, the temperature decreases and energy is taken in.
- B If the reaction is endothermic, the temperature increases and energy is given out.
- C If the reaction is exothermic, the temperature decreases and energy is given out.
- D If the reaction is exothermic, the temperature increases and energy is taken in.

13 Power stations produce electrical energy from different fuels.

Which fuel causes least pollution to the atmosphere?

- A coal
- B fuel oil
- C natural gas
- D radioactive isotopes

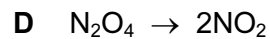
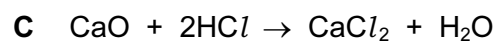
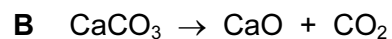
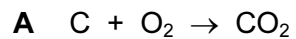
14 A student was investigating the reaction between marble chips and dilute hydrochloric acid.



Which changes would reduce the rate of reaction?

	temperature of acid	concentration of acid	surface area of marble chips
A	decrease	decrease	decrease
B	decrease	decrease	increase
C	increase	decrease	decrease
D	increase	increase	increase

15 Which equation shows an oxidation reaction?



16 In separate experiments, a catalyst is added to a reaction mixture and the temperature of the mixture is decreased.

What are the effects of these changes on the rate of the reaction?

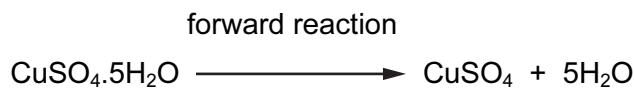
	catalyst added	temperature decreased
A	faster	faster
B	faster	slower
C	slower	faster
D	slower	slower

17 Different plants grow best under different pH conditions.

Which plant grows best in alkaline soil?

	plant	grows best in soil at pH
A	cabbage	6-8
B	potato	4-7
C	strawberry	5-7
D	wheat	6-7

- 18 The equation shows a reaction that is reversed by changing the conditions.



How can the forward reaction be reversed?

	by adding water	by heating
A	✓	✓
B	✓	x
C	x	✓
D	x	x

- 19 Element X forms an oxide, XO, that neutralises sulfuric acid.

Which row describes X and XO?

	element X	nature of oxide, XO
A	metal	acidic
B	metal	basic
C	non-metal	acidic
D	non-metal	basic

- 20 Copper carbonate reacts with dilute sulfuric acid to make copper sulfate.



Which row gives the correct order of steps for making copper sulfate crystals?

	step 1	step 2	step 3	step 4
A	add excess acid to the copper carbonate	filter	evaporate filtrate to point of crystallisation	leave to cool
B	add excess acid to the copper carbonate	filter	evaporate to dryness	leave to cool
C	add excess copper carbonate to the acid	evaporate to point of crystallisation	leave to cool	filter
D	add excess copper carbonate to the acid	filter	evaporate filtrate to point of crystallisation	leave to cool

21 Element X is a non-metal.

In which position of the Periodic Table could element X be found?

- A at the bottom of Group I
- B at the top of Group 0
- C at the top of Group I
- D in the transition elements

22 Aqueous sodium hydroxide is added to solid X and the mixture is heated.

A green precipitate is formed and an alkaline gas is given off.

Which ions are present in X?

- A NH_4^+ and Fe^{2+}
- B NH_4^+ and Fe^{3+}
- C OH^- and Fe^{2+}
- D OH^- and Fe^{3+}

23 A student carried out an experiment to find the order of reactivity of five metals. They were tested with cold water, hot water and steam and the results recorded in a table.

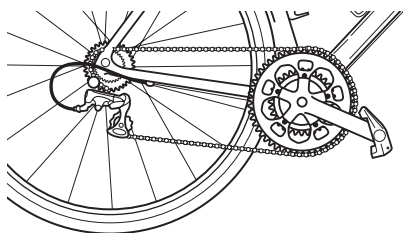
metal	cold water	hot water	steam
V	no reaction	reacts slowly	vigorous reaction
W	no reaction	no reaction	slow reaction
X	reacts slowly	vigorous reaction	not attempted
Y	no reaction	no reaction	no reaction
Z	vigorous reaction	explosive reaction	not attempted

What is the order of reactivity of these metals?

	most reactive		→	least reactive	
A	V	W		X	Z
B	W	X		V	Y
C	Z	X		W	Y
D	Z	X		W	V

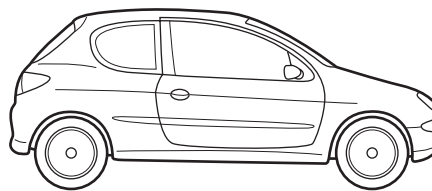
27 Which object is likely to be made from stainless steel?

A



bicycle chain

B



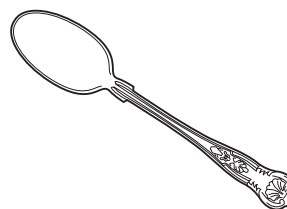
car body

C



can of beans

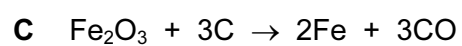
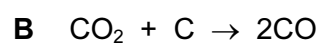
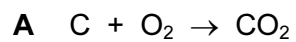
D



teaspoon

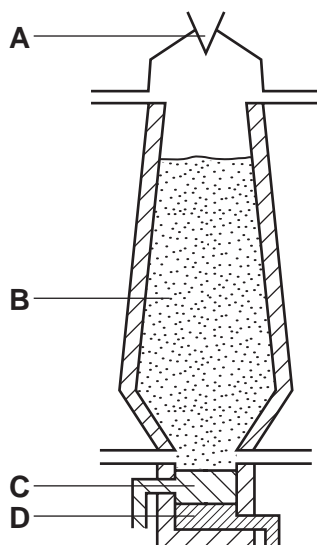
28 Four reactions that take place in the blast furnace to produce iron are shown.

Which reaction is used to keep the furnace hot?

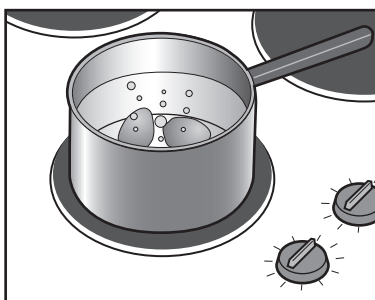


29 The diagram shows a blast furnace.

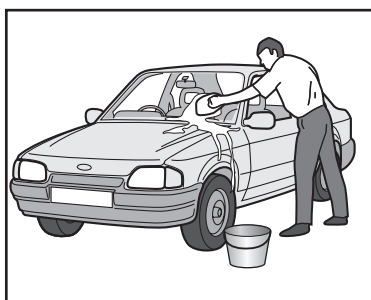
In which part is iron ore changed to iron?



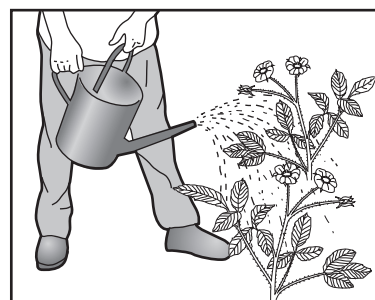
30 The diagram shows some uses of water in the home.



1



2



3

For which uses is it important for the water to have been treated?

- A** 1 only **B** 2 only **C** 3 only **D** 1, 2 and 3

31 A piece of uncoated iron and three pieces of iron with various coatings were left exposed to the air.

Which piece of iron would rust?

- A** the painted piece
B the tin-coated piece
C the uncoated piece
D the zinc-coated piece

32 Which compound would **not** be an effective fertiliser?

- A ammonium nitrate, NH_4NO_3
- B calcium oxide, CaO
- C calcium phosphate, $\text{Ca}_3(\text{PO}_4)_2$
- D potassium nitrate, KNO_3

33 Sulfur dioxide, SO_2 , nitrogen dioxide, NO_2 , and carbon monoxide, CO , are air pollutants.

Which row correctly shows their major source?

	motor car engines	power stations
A	CO	NO_2, SO_2
B	NO_2, CO	SO_2
C	SO_2, NO_2	CO
D	SO_2	NO_2, CO

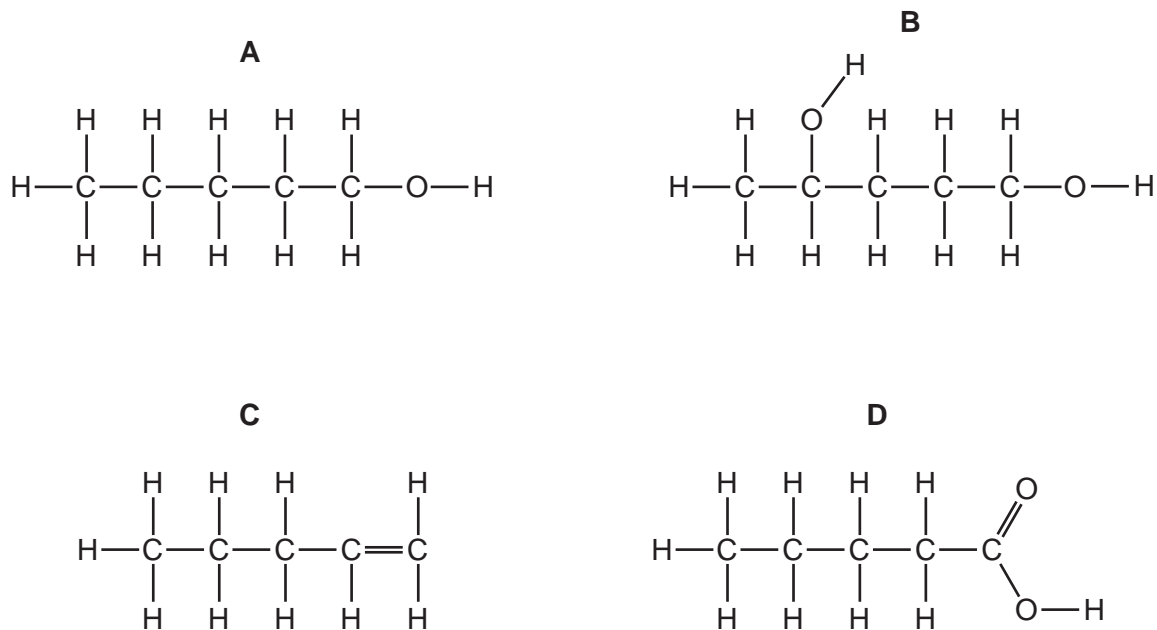
34 Which process does **not** produce carbon dioxide?

- A combustion of methane
- B fermentation of sugar
- C polymerisation of ethene
- D respiration

35 Which pollutant gas is produced by the decomposition of vegetation?

- A carbon monoxide
- B methane
- C nitrogen oxide
- D sulfur dioxide

36 Which diagram shows the structure of pentanoic acid?



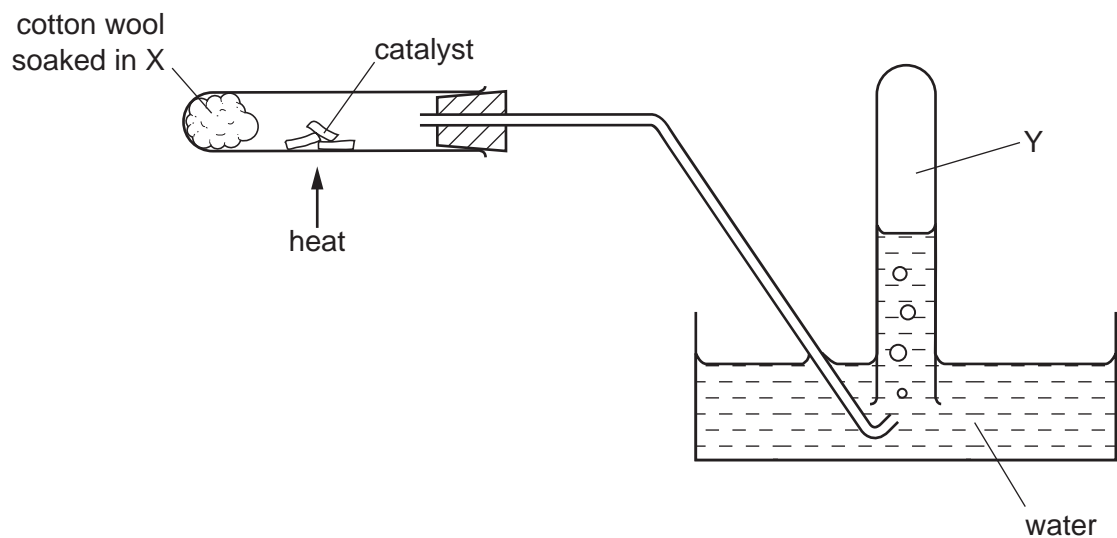
37 The table shows the composition of four different types of petroleum (crude oil).

fraction	Arabian Heavy /%	Arabian Light /%	Iranian Heavy /%	North Sea /%
gasoline	18	21	21	23
kerosene	11.5	13	13	15
diesel oil	18	20	20	24
fuel oil	52.5	46	46	38

Which type of petroleum is best for the motor vehicle industry?

- A** Arabian Heavy
- B** Arabian Light
- C** Iranian Heavy
- D** North Sea

38 The diagram shows the cracking of substance X.



Which type of organic compound is found in Y, which is **not** present in X?

- A acid
 B alcohol
 C alkane
 D alkene
- 39 In which reaction could one of the products belong to the same homologous series as the organic reactant?
- A addition of steam to ethene
 B combustion of an alkane
 C cracking of an alkane
 D polymerisation of ethene
- 40 Ethanol is produced from either ethene or sugar.

Which type of chemical reaction is used in each case?

	ethene → ethanol	sugar → ethanol
A	addition	fermentation
B	addition	fractional distillation
C	distillation	fermentation
D	distillation	fractional distillation

DATA SHEET
The Periodic Table of the Elements

		Group																		
		I	II	III	IV	V	VI	VII	VIII	IX	X									
		1 H Hydrogen 1																		
7	9	Li Lithium 3	Be Beryllium 4									He Helium 2								
23	24	Na Sodium 11	Mg Magnesium 12									Ne Neon 10								
39	40	K Potassium 19	Ca Calcium 20	45 Sc Scandium 21	48 Ti Titanium 22	51 V Vanadium 23	52 Cr Chromium 24	55 Mn Manganese 25	56 Fe Iron 26	59 Co Cobalt 27	59 Ni Nickel 28	64 Cu Copper 29	65 Zn Zinc 30	70 Ga Gallium 31	73 Ge Germanium 32	75 As Arsenic 33	79 Se Selenium 34	80 Br Bromine 35	84 Kr Krypton 36	
85	88	Rb Rubidium 37	Sr Strontium 38	89 Y Yttrium 39	91 Zr Zirconium 40	93 Nb Niobium 41	96 Mo Molybdenum 42	101 Ru Ruthenium 44	106 Pd Palladium 46	108 Ag Silver 47	112 Cd Cadmium 48	115 In Indium 49	119 Sn Tin 50	122 Sb Antimony 51	128 Te Tellurium 52	127 I Iodine 53	131 Xe Xenon 54			
133	137	Cs Caesium 55	Ba Barium 56	139 La Lanthanum 57	178 Hf Hafnium 72	181 Ta Tantalum 73	184 W Tungsten 74	190 Os Osmium 76	195 Pt Platinum 78	197 Au Gold 79	201 Hg Mercury 80	204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	210 Po Polonium 84	210 At Astatine 85	210 Rn Radon 86			
	226	Fr Francium 87	Ra Radium 88	227 Ac Actinium 89																
													*58-71 Lanthanoid series		†90-103 Actinoid series					
		a		X		a = relative atomic mass		X = atomic symbol		b = proton (atomic) number										
		b		Key																
		140 Ce Cerium 58	141 Pr Praseodymium 59	144 Nd Neodymium 60	146 Pm Promethium 61	150 Sm Samarium 62	152 Eu Europium 63	157 Gd Gadolinium 64	159 Tb Terbium 65	162 Dy Dysprosium 66	165 Ho Holmium 67	167 Er Erbium 68	169 Tm Thulium 69	173 Yb Ytterbium 70	175 Lu Lutetium 71					
		232 Th Thorium 90	238 Pa Protactinium 91	238 U Uranium 92	238 Np Neptunium 93	244 Pu Plutonium 94	244 Am Americium 95	244 Cm Curium 96	244 Bk Berkelium 97	244 Cf Californium 98	244 Es Einsteinium 99	244 Fm Fermium 100	244 Md Mendelevium 101	244 No Nobelium 102	244 Lr Lawrencium 103					

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

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